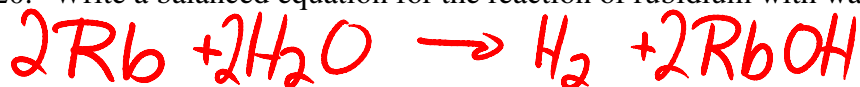


20. Write a balanced equation for the reaction of rubidium with water



21. Write a balanced equation for the reaction of lithium with air.



22. Which gas is used to fill ordinary electric light bulbs? *Argon* Why? *Noble Gas, will not support combustion of the filament like ordinary air would.*

23. Helium is the 2nd most abundant element in the universe?

24. In an **ionic** bond, electrons are transferred from a metal \rightarrow non-metal

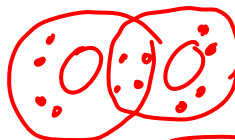
25. Give three examples of **ionic compounds**. NaCl, CaBr₂, AlI₃, NH₄I etc.

26. Ionic compounds have high melting points. From this we know they have (strong/weak) strong bonds between the ions.

27. Bonds in which electrons are shared equally or almost equally are called covalent (non polar covalent)

28. Draw the electron-dot structure of a **diatomic oxygen molecule**

A **diatomic iodine molecule**



A **diatomic nitrogen molecule**

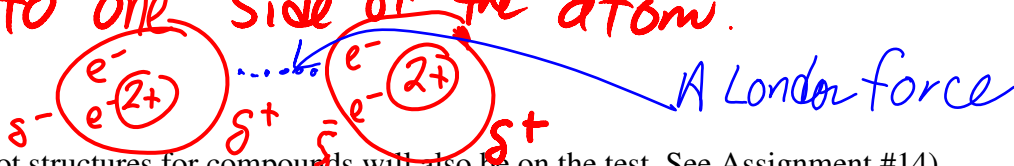


29. What is meant by a **polar covalent** bond?

A bond in which electrons are shared unequally between 2 atoms

30. Explain what causes **London forces**. Use a diagram

- temporary dipoles resulting from e⁻s being more to one side of the atom.



(NOTE: Electron-dot structures for compounds will also be on the test. See Assignment #14)